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ARRANGEMENT OF
ELECTRONS IN
ATOMS93 FIGURE

4-3The photoelectric effect: electromagnetic radiation strikes the surface of the metal, ejecting electrons from the metal and creating an electric current.

Frequency and wavelength are mathematically related

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to each other. For electromagnetic radiation, this relationship is written as follows.

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SECTION 3 SHORT

ANSWER Answer the following questions in the space provided. 1.

State the Pauli exclusion principle, and use it to explain why electrons in the same orbital must have opposite spin states.

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The Pauli exclusion principle states that no two electrons in an atom may have the

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Chapter Four

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Periodic Law] Chapter
Six [Chemical Bonding]

... Arrangement of
Electrons. Interactives:
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Taken from the book Modern Chemistry by Holt, Rinehart, and Winston on Chapters 4 and 5, which deals with

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...

Arrangement of the
Electrons Chapter 4
(Electron

Configurations)

Electron Behavior. ...

-ordered arrangement
by wavelength or
frequency for all forms
of electromagnetic
radiation. Parts of the
wave, Wavelength-

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lambda (λ) The

distance between
corresponding points
on adjacent waves.

Units: m, nm, cm, or Å

Arrangement of the Electrons Chapter 4

Modern Chemistry 1

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CHAPTER 4 REVIEW

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Teacher Notes and

Answers Chapter 4

SECTION 1 SHORT

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Chapter 4

ANSWER 1. In order for an electron to be ejected from a metal surface, the electron must be struck by a single photon with at least the minimum energy needed to knock the electron loose. 2.

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4-1 CHEMISTRY

CHAPTER 4

(Arrangement of
Electrons) The lowest
energy state of an
atom is its ground
state. (usually it's the

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lowest levels) A photon is a particle of electromagnetic radiation having zero mass and carrying a quantum of energy. When a photon strikes an atom it gives the atom more energy. If enough photons strike an atom it may

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Review CHAPTER 4:
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Include graphic organizer (s) for this chapter The following pages contain the bulk (but not all) of the information for the chapter 4 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc.

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Development ...

related to the behavior of electrons. CHAPTER 4 Neon Walkway.

ARRANGEMENT OF ELECTRONS IN ATOMS
97 SECTION 1 O

OBJECTIVES Explain the mathematical relationship among the speed, wavelength, and frequency of electromagnetic radiation. Discuss the dual wave-particle nature of light.

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and Chemical
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Chapter 4 Review
Answers ...**

- _____ 4. How many electrons are present in an atom of calcium that has the electron configuration $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$?
- a. 6 b. 16 c. 20 d. 36
- _____ 5. The ground-state electron

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configuration of neon is $1s^2 2s^2 2p^6$. In this arrangement, how many of neon's p orbitals are completely filled? a. 1 b. 2 c. 3 d. 6

Name Class Date

Arrangement of ...

Assessment

Arrangement of Electrons in Atoms

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Teacher Notes and
Answers Chapter 4

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SECTION 1 SHORT ANSWER 1. In order for an electron to be ejected from a metal surface, the electron must be struck by a single photon with at least the minimum energy needed to knock the electron loose. 2. The ground state is the lowest energy state of the atom.

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Section 2 Answers

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...

Chapter 4, Lesson 3:
The Periodic Table and
Energy-Level Models.
Key Concepts. • The
electrons surrounding
an atom are located in
regions around the
nucleus called “energy
levels”. • An energy
level represents the

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3-dimensional space surrounding the nucleus where electrons are most likely to be.

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